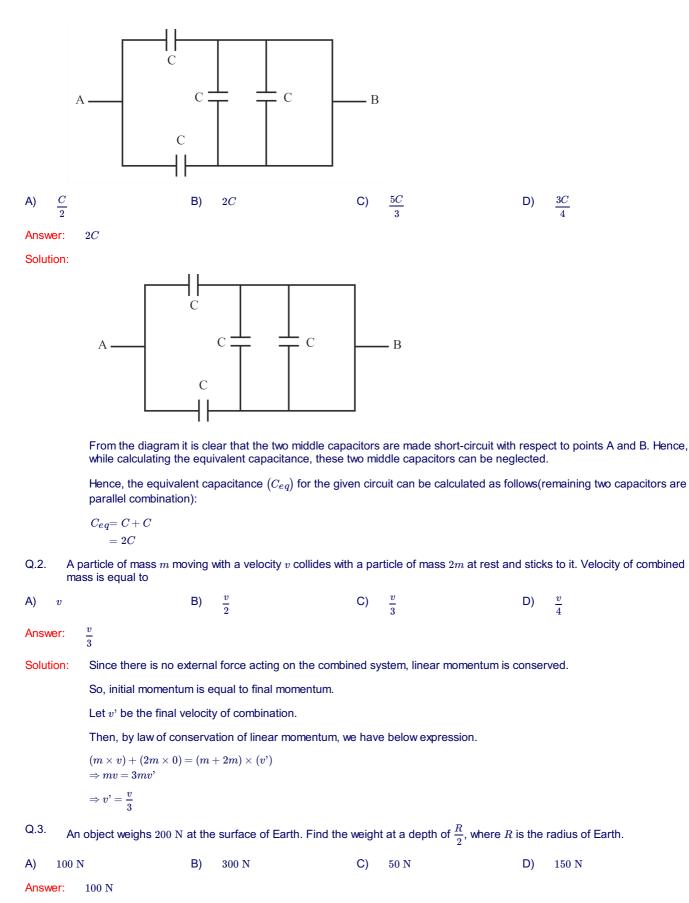


JEE Main 2023 (Session 2)

April 10 Shift 1



Physics



Q.1. Find the equivalent capacitance across points A and B in the given electrical circuit.



Solution: The variation of acceleration due to gravity with depth can be expressed as follows:

$$g' = g\left(1 - \frac{d}{R}\right) \dots \left(1\right)$$

Substitute $\frac{R}{2}$ for d into equation (1) to obtain the new acceleration due to gravity at the given depth.

$$g' = g\left(1 - \frac{\frac{R}{2}}{R}\right)$$
$$= \frac{g}{2}$$

Since, the acceleration due to gravity is halved, the weight of the object will also be halved at the given depth.

Hence, the weight of the object at the given depth is 100 N.

Q.4. For an object radiating heat at 300 K, the wavelength corresponding to maximum intensity is λ . If the temperature of body is increased by 300 K, the new wavelength corresponding to maximum intensity will be

A)
$$\frac{\lambda}{2}$$
 B) 2λ C) λ D) $\frac{5\lambda}{2}$

Answer:

 $\frac{\lambda}{2}$

Solution: According to Wien's displacement law,

 $\lambda_m T = \text{constant}$

where, λ_m is wavelength of maximum intensity and T is absolute temperature.

Let λ_m ' be the required wavelength at 600 K (Since temperature is increased by 300 K)

$$\Rightarrow \lambda_m' imes 600 = \lambda imes 300$$

 $\Rightarrow \lambda_m' = rac{\lambda}{2}$

Q.5. A particle, when projected at 15^o with horizontal, has a range of 50 m. Find the range when projected at 45^o with horizontal.

50 m B) 100 m C) 80 m D) 120 m

Answer: 100 m

Solution:

A)

The formula for horizontal range of a projectile is given by, $R = rac{u^2 \sin(2 heta)}{g}$

For angle of projection, $\theta = 15^{o}$, range is given by,

$$R = \frac{u^2 \sin(30^o)}{g}$$
$$= \frac{u^2}{2g} = 50 \text{ m } \left(\text{Given}\right)$$

For angle of projection, $\theta = 45^{\circ}$, range is given by,

$$R' = \frac{u^2 \sin(90^o)}{g}$$
$$= \frac{u^2}{g} = 100 \text{ m}$$

Q.6. Statement (1) : An LCR circuit connected to an AC source has maximum average power at resonance.

Statement (2): A resistor only circuit with zero phase difference has maximum average power.

A) (1) and (2) both are correct

B) (1) is correct but (2) is incorrect

C) (1) is incorrect but (2) is correct

D) (1) and (2) both are incorrect

Answer: (1) and (2) both are correct



Solution: We know that at resonance, an LCR circuit connected to an AC source has current amplitude at its maximum value.

In addition, average power reaches a maximum when impedance (*Z*), which depends on the frequency, is a minimum, that is, when $X_C = X_L$ and Z = R. Thus, at resonance, the average power output of the source in an LCR series circuit is a maximum.

Thus, both the statements are correct.

- Q.7. A monoatomic gas initially at pressure *P* and volume *V* is compressed to $\frac{1}{8}$ of its volume adiabatically. Final pressure of the gas is equal to
- A) 4P B) 8P C) 16P D) 32P

Answer: 32P

Solution: The initial and final parameters of the gas are related by the equation

$$P_i V_i^{\gamma} = P_f V_f^{\gamma}$$
$$P_f = P_i \left(\frac{V_i}{V_f}\right)^{\gamma} \dots \left(1\right)$$

Substitute the values of the known parameters into equation (1) to obtain the final pressure of the gas.

$$P_{f} = P\left(\frac{V_{i}}{\frac{V_{i}}{8}}\right)^{\frac{5}{3}}$$
$$= 32P$$

Q.8. What is the maximum percentage error in the measurement of quantity *l*, if it is given by $l = \frac{a^2b^3}{c\sqrt{d}}$? Given the percentage error in the calculation of *a*, *b*, *c* and *d* are 1%, 2%, 3% and 4% respectively.

A) 4% B) 12% C) 9% D) 13%

Answer: 13%

Solution: The formula to calculate the percentage error in measuring the quantity *l* can be expressed as

$$\frac{\Delta l}{l}\% = \left[\left(2\frac{\Delta a}{a} + 3\frac{\Delta b}{b} + \frac{\Delta c}{c} + \frac{1}{2}\frac{\Delta d}{d} \right) \times 100 \right]\% \quad \dots \left(1 \frac{\Delta a}{b} + \frac{\Delta c}{b} + \frac{1}{2}\frac{\Delta d}{d} \right) \times 100 \right]\%$$

Substitute the values of the known parameters into equation (1) to calculate the required percentage error to measure the given quantity.

$$\frac{\Delta l}{l} \% = \left[\left(2 \times 0.01 + 3 \times 0.02 + 0.03 + \frac{1}{2} \times 0.04 \right) \times 100 \right] \%$$

= 13%

Q.9. The equation of progressive wave is $y = 5 \sin(6t + 0.03x)$. Find the speed of wave.

Answer: 200

Solution: Here, the direction of wave movement is along -X axis.

The standard equation of a wave moving in -X direction is given by,

 $y = A \sin \left(\omega t + k x \right)$

where, the speed of wave is given by, speed = $\frac{\omega}{k}$

Upon substitution, we get

speed =
$$\frac{6}{0.03}$$
$$= 200 \text{ m s}^{-1}$$

Q.10. Earth shrinks to $\frac{1}{64}$ times of its initial volume. Time period of Earth rotation is found to be $\frac{24}{x}$ hr. Find the value of x.



Solution: As the radius of Earth is decreased, its moment of inertia will also change.

The time period (T) and the angular frequency (ω) are related by the relation

$$\omega = \frac{2\pi}{T} \dots (1)$$

Considering Earth as a perfect sphere of radius R, its initial angular momentum (L) is given by

$$L = \left(\frac{2}{5}MR^2\right)\omega \quad \dots \left(2\right)$$

From equations (1) and (2), it implies that

$$L = \frac{2}{5}MR^2 \times \frac{2\pi}{T}$$
$$\Rightarrow T = \frac{4\pi}{5L}MR^2 \quad \dots (3)$$

The radius and volume of a spere is related by the relation

$$R \propto V^{\frac{1}{3}} \dots \left(4\right)$$

Using equation (4), the new radius of the Earth (R') is given by

$$\frac{R'}{R} = \left(\frac{\frac{V}{64}}{V}\right)^{\frac{1}{3}}$$
$$\Rightarrow R' = \frac{R}{4} \dots (5)$$

Substitute the expression for the new radius into equation (3) to obtain the new time period (T') when the volume is decreased.

$$T' = \frac{4\pi}{5L} M\left(\frac{R}{4}\right)^2$$
$$= \frac{T}{16} \dots (6)$$

Comparing equation (6) with the given expression for the new time period, it can be concluded that x = 16.

Q.11. A conducting rod of length 1 m is moved across a magnetic field of 0.15 T, with constant speed of 4 m s^{-1} . Find force (in N) on rod.

Answer:

0

Solution: The formula to calculate the emf (ε) induced in the moving rod of length *l*, moving with a speed of *v* under a magnetic field *B* can be expressed as(in this case length of the rod is perpendicular to both magnetic field and velocity of the rod)

$$\varepsilon = Blv \dots (1)$$

Substitute the values of the known parameters into equation (1) to calculate the required induced emf.

$$arepsilon = 0.15 \mathrm{~T} imes 1 \mathrm{~m} imes 4 \mathrm{~m} \mathrm{~s}^{-1}$$

= 0.6 V

Now as the rod is not connected to any circuit, there will not be any current flowing & therefore no magnetic force on the rod.

Hence, the required value of the force is 0.

Q.12. 10 resistors each of 10 Ω resistance when connected together give minimum equivalent resistance R_1 and maximum equivalent resistance R_2 among various possible combinations. So $\frac{R_2}{R_1}$ is equal to



A set of resistors connected in a circuit produces maximum value of equivalent resistance when they are connected in series Solution: combination. On the other hand, the equivalent resistance is minimum when the set of resistors are connected in parallel combination.

When ten resistors of resistance 10 Ω each are connected in series combination, the equivalent resistance (R_2) is given by

 $R_2 = 10 \ \Omega + 10 \ \Omega + \dots \dots 10 \ \text{times}$ $= 100 \Omega \ldots (1)$

When these are connected in parallel combination, the equivalent resistance (R_1) of the circuit is given by

 $\frac{1}{R_1} = \frac{1}{10 \Omega} + \frac{1}{10 \Omega} + \dots \dots 10$ times $\Rightarrow R_1 = 1 \Omega \dots (2)$

Divide equation (1) by equation (2) to calculate the required ratio.

 $\frac{R_2}{R_1} = \frac{100 \,\Omega}{1 \,\Omega}$ = 100

Q.13. In an AM wave, amplitude of modulating wave = 3 units and amplitude of carrier wave = 15 units. Find the ratio of maximum to minimum intensity $\frac{Imax}{I_{min}}$

Answer: 2.25

Solution: Given.

Amplitude of modulating wave, $(A_M) = 3$ units

Amplitude of carrier wave, $(A_C) = 15$ units

As both modulating and carrier waves are superimposed, $A_{max} = A_C + A_M$

and $A_{min} = A_C - A_M$

Since intensity is proportional to the square of amplitude.

$$\begin{aligned} \frac{Imax}{Imin} &= \left(\frac{AC^{+}AM}{AC^{-}AM}\right)^2 \\ &= \left(\frac{15+3}{15-3}\right)^2 \\ &= \left(\frac{18}{12}\right)^2 \\ &= 1.5^2 = 2.25 \end{aligned}$$

Chemistry

Q.14.

- Stabiliser used for concentration of sulphide ore is A) Fatty acids B) Pine iol C) Cresols D) **Xanthates**
- Answer: Cresols
- Froth flotation method is used to remove gangue from the sulphide ores by the formation of a powdered ores' suspension in Solution: water with the use of collectors and stabilisers. The role of the stabiliser in the froth flotation process is used to stabilise the froth such as cresols and aniline. It increases the non-wettability of the mineral particles.
- Which of the following complex compound is diamagnetic and low spin? Q.15.

A) [Co	$(NH_3)_6]^{+3}$	B)	$[{ m CoCl}_6]^{-3}$	C)	$[{ m CoF}_6]^{-3}$	D)	$\left[\mathrm{Fe}\left(\mathrm{H_{2}O}\right)_{6} ight]^{+3}$
Answer:	$\left[\mathrm{Co}(\mathrm{NH}_3)_6\right]^{+3}$						

Solution: $\ln \left[\operatorname{Co} \left(\mathrm{NH}_3 \right)_6 \right]^{+3}$ the oxidation state of cobalt is +3. Ammonia is a strong field ligand due to that pairing of electrons takes place. The hybridisation of central metal atom is d^2sp^3 . Since it has no unpaired electrons, it is diamagnetic.

> Strong field ligands cause larger splitting of d orbitals and pairing of electrons is favoured. Hence, it is a low spin complex.



A O ⊂ O linkage B) S ≤ S linkage C) Vanderwalis force D) Hydrogen bonding Answer: O − O linkage B) S ≤ S linkage C) Vanderwalis force D) Hydrogen bonding Answer: O − O linkage B) S ≤ S linkage C) Vanderwalis force D) Hydrogen bonding Answer: In accordary structure of polypeptides, the peptide chains are stabilised by hydrogen bonding. and hydrophobic interactions, and the stabilised by various types of bonds, including disulphide bonds, ionic bonds, hydrogen bonds, and hydrophobic interactions, among others. So among the given options O · O linkage doesn't stabilise secondary and tertiary protein. C1.1 Which compound does not oxist from the following? D) (NIT_d)2BcF4 Answer: NaO2 C) PbE14 D) (NIT_d)2BcF4 Answer: NaO2 Solution: Anong the alkali metals, south more otors of more other other other atkali metals (such as polassium, rubidum, the oceal mark in thom to form supervades. With excess nowage. D = Solution: D = Solution: Answer: DoXdation B) Roduction C) hydroytes in the assolution to form supervade ion (O2 ·) under normal condinons. Therefore, the compound that does not exist from the given options is NaO2. <	Q.16. One which does not st	tabilise secondary and ter	tiarv protein.			
Answer: 0 - 0 linkage Solution: In primary amines, peptide linkages (alse known as peptide brotels) are formed between the amine oroup of one amine acid, and the carboxy groups of adjacent amine acid, exaulting in the formation of a peptide chain. In secondary structure of polypeptides, the peptide brotes are stabilised by hydrogen bonding between the amine acid. In trainay structure, the protein chains is folded into a three-dimensional conformation, and the stabilized by various types of bonds, including disulphide bonds, ionic bonds, hydrogen bonds, and hydrophobic interactions, among others. So among the given options O = O linkage doesn't stabilise secondary and tertiary protein. 0.17. Which compound does not exist from the following? A) BxClig D) NBAC C) PhEt4 D) Nareser: No2 Solution: Among the alkali metals, sodium is known to form peroxides, whereas the other alkali metals (such as potassium, rubidium, and desium) meroxide (Nx20). However, sodium does not form the superiods (Nx20) as well as with excess orgoning the form south meroxide (Nx20). However, sodium does not form the superiods (Nx20) as well as with excess orgoning the forma superiods. Solution does not form the superiods (Nx20). 0.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent D) Breaking 10. Solution: The reaction is as follows, 2PrSO ₄ → Prc ₂ O ₃ + SO ₂ + SO ₃ D) Breaking		-			D)	Hydrogen bonding
and the carbox/ group of another amino acid, resulting in the formation of a peptide chain. In secondary structure of polypeptides, the peptide chains are stabilised by hydrogen bonding between the amino and carbox/ groups of adjacent amino acids. In tertiary structure, the protein chain is folded into a three-dimensional conformation, and the stabilised by various types of bonds, including disulphide bonds, lonic bonds, hydrogen bonds, and hydrophobic interactions, among others. So among the given options 0 – 0 linkage doesn't stabilise secondary and tertiary protein. Q.17. Which compound does not exist from the following? A) ReCl ₂ B) No ₂ C) PhEt ₄ D) (NH ₄) ₂ BeF ₄ Answer: Na ₂ Solution: Among the alkali metals, sodium is known to form peroxdes, whereas the other alkalise (such as potassium, rubidium, and deciming are known to form superoxdes. Sodium reacts with oxygen to form sodium code (Na ₂ O) se well as with exession codes in the structure, the compound that does not terms the superoxde ion (O ₂ ⁻⁾ under normal conditions. Therefore, the compound that does not terms the superoxde ion (O ₂ ⁻⁾ under normal conditions. Therefore, the compound that does not form the superoxde ion (O ₂ ⁻⁾ under normal conditions. Therefore, the compound that does not terms sulphate (FaSO ₄) is formed. On prolonged heating Fierous ion oxidiated to ferric lons. The reaction is as follows, 2Prolong sted heating, ferrous sulphate crystals lose water and aniyorus ferrous sulphate (FaSO ₄) is formed. On prolonged heating fierous ion oxidiated to ferric lons. The reaction is as follows, 2Prolo ₄ + Fe ₂ O ₃ + SO ₂ + SO ₂ C.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive. Positive B) Positive, Positive B, Positive, Positive Co.20. Read the following to statements Statement I: Socret, and II are correct. A) Solution: The refore, the enthalpy of adsorption in negative. Statement I: Socret, or is more solution or micell a subject substitue is pos	Answer: 0 – 0 linkage	, .				
carboxyl groups of adjacent amino acids. In tertiary structure, the protein chain is folded into a three-dimensional conformation, and the stabilised by various types of books, including is onic books, hydrogen bonds, and hydrophobic interactions, among others. So among the given options 0 – 0 linkage doesn't stabilise secondary and tertiary protein. Q.17. Which compound does not exist from the following? A) BcCl2 B) NaO2 C) PbEt4 D) (NII4)2BcF4 Answer: NaO2 Solution: Among the alkali metals, sodium is known to form peroxides, whereas the other alkali metals (such as potassium, rubidium, rubidium, rubidium, and the submovide (Nay 0) as well as with excess oxygen to form sodium peroxide (Nay 02). However, sodium does not form the superoxide ino (O2 ⁻) under normal continuon. Therefore, the compound that does not exist from the given options is NaO2. Q.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent A) A) Oxidation B) Reduction C) hydrolysis D) Breaking Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (FeSO ₄) is formed. On prolonged heating ferrous in oxidied to ferric ions. The reaction is as follows, 2FeSO ₄ → Fe2O ₃ + SO ₂ + SO ₃ Q.19. Enthalpy of adsorption and enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, entha						
binds, including disulphide bonds, ionic bonds, hydrogen bonds, and hydrophobic interactions, among others: So among the given options O − 0 linkage doesn't stabilise secondary and tortiary protein. Q.17. Which compound does not exist from the following? B) NaO2 C) PhEt4 D) (NH4j2BaF4 Answer: NaO2 B) NaO2 C) PhEt4 D) (NH4j2BaF4 Answer: NaO2 Solution: Among the alkali metals, sodium is known to form peroxides, whereas the other alkali metals (such as potassium, rubidium, and cesium) are known to form superoxides. Sodium reacts with oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium peroxide (Na ₂ O). Howwere: sodium does not exist from the given options is NaO2. Q.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent D) B Reaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking C19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively D) N positive. Positive B) Positive, Negative and enthalpy of adsorption. In the vast mejority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption. In the vast mejority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption. In the vast mejority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption. In the vast mejority of cases, adsorption is an exothermic in nature. Therefore, the			eptide chain	is are stabilised by hydro	gen bono	ding between the amino and
Q. 17. Which compound does not exist from the following? A) BcCl2 B) NaO2 C) PbEt4 D) (NII J)2BeF4 Answer: NaO2 Solution: Among the alkali metals, sodium is known to form peroxides, whereas the other alkali metals (such as potassium, rubidium, and cesium) are known to form superoxides. Sodium reads with oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium oxide. Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction of micelle are respectively A) No positive, Positive B) Positive, Negative D) Negative, Negative A) Positive, Positive B) Positive, Negative, Positive D) Negative, Positive Negative, Positive D) Negative, Positive Negative, Positive Negative, Positive, Negative, Positive, Negative, Positive <td></td> <td></td> <td></td> <td></td> <td></td> <td></td>						
A) BcCl2 B) NaO2 C) PbEt4 D) (NH4)2BeF4 Answer: NaO2 Solution: Among the alkali metals, sodium is known to form peroxides, whereas the other alkali metals (such as potassium, rubidium, and desium) are known to form superoxide (Na)O2). However, sodium does not form the superoxide ion (O2 ⁻¹) under normal conditions. Therefore, the compound that does not exist from the given options is NaO2. C.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent A) Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction of incelle are respectively Answer: Answer: The reaction is as follows, adortive. Negative D) Negative. Negative 2.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) No. Negative, Positive D) Negative. Negative Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbate is called enthalpy of dasorption. In the vast majority of cases, adsorbed is an exothermic in nature. Therefore, enclabargot adsorption is neansubterwic in mature. Therefore, enclabargot adsorptio	So among the give	n options $\mathrm{O}-\mathrm{O}$ linkage d	oesn't stabili	ise secondary and tertiar	y protein	
Answer: NaO2 Solution: Among the alkali metals, sodium is known to form peroxides. Sodium reads with oxygen to form sodium oxde (Na ₂ O) as well as with excess oxygen to form sodium oxde (Na ₂ O) as well as with excess oxygen to form sodium peroxide (Na ₂ O). However, sodium does not form the gueen potions is NaO2. Q.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent A) Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (FeSO ₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, 2FeSO ₄ → Fe ₂ O ₃ + SO ₂ + SO ₃ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative C) Negative, Positive Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbation is nexothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means All > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the	Q.17. Which compound does	s not exist from the followi	ng?			
Solution: Another alkali metals, sodium is known to form peroxides. Sodium reads with oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium oxide (Na ₂ O) as well as with excess oxygen to form sodium peroxide ion (O ₂ ⁻) under normal conditions. Therefore, the compound that does not exist from the given options is NaO2. Q.18. Prolongated heating of ferrous ammonium subplate is avoided to prevent A) Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation B) Reduction C) hydrolysis D) Breaking Solution: On heating, ferrous subplate crystals lose water and anhydrous ferrous sulphate (FeSO ₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, 2FeSO ₄ → Fe ₂ O ₃ + SO ₂ + SO ₃ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively D) Negative, Negative Solution: Reserve: Negative, Positive D) Negative, Negative Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, entimalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means L	A) $BeCl_2$	B) NaO ₂	C)	${\rm PbEt}_4$	D)	$(\mathrm{NH}_4)_2\mathrm{BeF}_4$
and cesium) are known to form superoxides. Sodium reads with oxygen to form souperoxide (Na ₂ O) is well as with excess oxygen to form solum oxide (Na ₂ O) is well as with excess oxygen to form solum oxide (Na ₂ O) is well as with excess oxygen to form solut on or normal conditions. Therefore, the compound that does not exist from the given options is Na ₀ ₂ . Q.18. Prolongated heating of ferrous ammonium sulphate is avoided to prevent A) Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (FeSO ₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, 2FeSO ₄ → Fe ₂ O ₃ + SO ₂ + SO ₃ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative C) Negative, Positive D) Negative, Negative Answer: Negative, Positive Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Notestioned is used in volumetric analysis. Statement I: Solution are out and II are correct. A) Both statements I and II are correct and II is incorrect. C) Statement I is correct and II is incorrect. Solution: Potassium dichromate is used in volumetric in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na ₂ Cr ₂ O ₇ is hygroscopic in nature and is more soluble in water than K ₂ Cr ₂ O ₇ .	Answer: NaO ₂					
A) Oxidation B) Reduction C) hydrolysis D) Breaking Answer: Oxidation Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (PeSO₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, 2FeSO₄ → Fe₂O₃ + SO₂ + SO₃ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative. D) Negative, Negative Answer: Negative, Positive B) Positive, Negative. D) Negative, Negative. Solution: The arount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbate is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: K₂Cr₂O₂ is more soluble in water than Na₂Cr₂O₁ A) A) Both statements I and II are correct. B) Both statements I and II are incorrect. C) C) Statement I is correct and II is incorrect. D)	and cesium) are kr oxygen to form sod	hown to form superoxides. lium peroxide (Na_2O_2) . He	Sodium read	cts with oxygen to form so um does not form the sup	odium oxi eroxide i	de (Na_2O) as well as with excess
Answer: Oxidation Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (FeSO ₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, $2FeSO_4 \rightarrow Fe_2O_3 + SO_2 + SO_3$ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: Rc ₂ Or ₂ O ₇ is more soluble in water than Na ₂ Cr ₂ O ₇ A) Both statements I and II are correct. D) C) Statement I is correct and II is incorrect. D) Statement I is correct and II is incorrect. C) Statement I is correct and II is incorrect. D) Statement I is correct and II is incorrect. C) Statement I is correct and II is incorrect. C) Statement I is correct and II is incorrect. D) Statement I is correct and	Q.18. Prolongated heating o	of ferrous ammonium sulpl	nate is avoid	ed to prevent		
Solution: On heating, ferrous sulphate crystals lose water and anhydrous ferrous sulphate (FeSO₄) is formed. On prolonged heating ferrous ion oxidised to ferric ions. The reaction is as follows, 2FeSO₄ → Fe2O₃ + SO₂ + SO₃ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative D) Negative, Negative, Negative Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: K2Ct2O7 is more soluble in water than Na2Ct2O7 A) Both statement I and II are correct. B) Both statement I is correct. C) Statement I is correct and II is incorrect. D) Statement I is correct. A: Both statement I is correct and II is incorrect. D) Statement I is correct. C) Statement I is correct and II is incorrect. D) Statement I is correct.	A) Oxidation	B) Reduction	C)	hydrolysis	D)	Breaking
ferrous ion oxidised to ferric ions. The reaction is as follows, $2FeSO_4 \rightarrow Fe_2O_3 + SO_2 + SO_3$ Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative D) Negative, Negative Answer: Negative, Positive B) Positive, Negative C) Negative, Positive D) Negative, Negative Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption is negative. Micelle formation decreases the stability of the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: Ry ₂ Cr ₂ O ₇ is more soluble in water than Na ₂ Cr ₂ O ₇ . A) Both statement I is correct and II is incorrect. D) Statement I is incorrect. C) Statement I is correct and II is incorrect. D) Statement I is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na ₂ Cr ₂ O ₇	Answer: Oxidation					
$\begin{array}{c c c c c c c c c c c c c c c c c c c $			ater and anh	ydrous ferrous sulphate	(FeSO4) i	s formed. On prolonged heating
 Q.19. Enthalpy of adsorption and enthalpy of formation of micelle are respectively A) Positive, Positive B) Positive, Negative C) Negative, Positive D) Negative, Negative Answer: Negative, Positive Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: K2Cr2O7 is more soluble in water than Na2Cr2O7 A) Both statements I and II are correct. B) Both statements I and II are incorrect. C) Statement I is correct and II is incorrect. D) Statement I is correct and II is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na₂Cr₂O₇: 	The reaction is as t	follows,				
 A) Positive, Positive B) Positive, Negative C) Negative, Positive D) Negative, Negative Answer: Negative, Positive Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement II: K₂Cr₂O₇ is more soluble in water than Na₂Cr₂O₇. A) Both statements I and II are correct. C) Statement I is correct and II is incorrect. C) Statement I is correct and II is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na₂Cr₂O₇ is hygroscopic in nature and is more soluble in water than K₂Cr₂O₇. 	$2\mathrm{FeSO}_4 ightarrow \mathrm{Fe}_2\mathrm{O}_3$ +	+ SO ₂ + SO ₃				
Answer: Negative, Positive Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means $\Delta H > 0$. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: Potassium dichromate is used in volumetric analysis. Statement I: RgCr207 is more soluble in water than Na2Cr207 A) Both statements I and II are correct. B) Both statements I and II are correct. C) Statement I is correct and II is incorrect. D) Statement I is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na ₂ Cr ₂ O ₇ is hygroscopic in nature and is more soluble in water than K ₂ Cr ₂ O ₇ .	Q.19. Enthalpy of adsorption	n and enthalpy of formatio	n of micelle a	are respectively		
Solution: The amount of heat evolved when one mole of an adsorbate (gas or liquid) is adsorbed on the surface of an adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative. Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement II: K2 Cr2 O7 is more soluble in water than Na2 Cr2 O7 A) Both statements I and II are correct. B) Both statements I and II are incorrect. C) Statement I is correct and II is incorrect. D) Statement I is incorrect. C) Statement I is correct and II is incorrect. D) Statement I is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na ₂ Cr ₂ O ₇ is hygroscopic in nature and is more soluble in water than K ₂ Cr ₂ O ₇ .	A) Positive, Positive	B) Positive, Negative	C)	Negative, Positive	D)	Negative, Negative
adsorbent is called enthalpy of adsorption. In the vast majority of cases, adsorption is an exothermic in nature. Therefore, enthalpy of adsorption is negative.Micelle formation decreases the stability of the colloidal solution so energy of the mixture should increase which means $\Delta H > 0$. Therefore, the enthalpy of formation of micelle is positive.Q.20.Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement II: $K_2 Cr_2 O_7$ is more soluble in water than $Na_2 Cr_2 O_7$ A)Both statements I and II are correct.B)B)Both statements I and II are correct.C)Statement I is correct and II is incorrect.C)Statement I is correct and II is incorrect.C)Statement I is correct and II is incorrect.Solution:Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. $Na_2 Cr_2 O_7$ is hygroscopic in nature and is more soluble in water than $K_2 Cr_2 O_7$.	Answer: Negative, Positive					
 means ΔH > 0. Therefore, the enthalpy of formation of micelle is positive. Q.20. Read the following two statements. Statement I: Potassium dichromate is used in volumetric analysis. Statement II: K₂Cr₂O₇ is more soluble in water than Na₂Cr₂O₇ A) Both statements I and II are correct. B) Both statements I and II are incorrect. C) Statement I is correct and II is incorrect. D) Statement I is correct and II is incorrect. Answer: Statement I is correct and II is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na₂Cr₂O₇ is hygroscopic in nature and is more soluble in water than K₂Cr₂O₇. 	adsorbent is	called enthalpy of adsorp	tion. In the v			
Statement I: Potassium dichromate is used in volumetric analysis. Statement II: K ₂ Cr ₂ O ₇ is more soluble in water than Na ₂ Cr ₂ O ₇ A) Both statements I and II are correct. B) Both statements I and II are incorrect. C) Statement I is correct and II is incorrect. D) Statement I is incorrect and II is correct. Answer: Statement I is correct and II is incorrect. D) Statement I is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na ₂ Cr ₂ O ₇ is hygroscopic in nature and is more soluble in water than K ₂ Cr ₂ O ₇ .					of the mix	dure should increase which
 C) Statement I is correct and II is incorrect. D) Statement I is incorrect and II is correct. Answer: Statement I is correct and II is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na₂Cr₂O₇is hygroscopic in nature and is more soluble in water than K₂Cr₂O₇. 	Statement I: Potassiun	m dichromate is used in vo				
Answer: Statement I is correct and II is incorrect. Solution: Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. Na2 Cr2 O7 is hygroscopic in nature and is more soluble in water than K2 Cr2 O7.	A) Both statements I and II a	re correct.	B)	Both statements I and	Il are inc	orrect.
Solution:Potassium dichromate is preferred over sodium dichromate in volumetric analysis, because sodium dichromate is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. $Na_2Cr_2O_7$ is hygroscopic in nature and is more soluble in water than $K_2Cr_2O_7$.	C) Statement I is correct and	d II is incorrect.	D)	Statement I is incorrec	ct and II is	s correct.
is hygroscopic in nature and, therefore, accurate weighing is not possible in normal atmospheric conditions. $Na_2Cr_2O_7$ is hygroscopic in nature and is more soluble in water than $K_2Cr_2O_7$.	Answer: Statement I is correct	ct and II is incorrect.				
Therefore, option C is correct.	$\operatorname{Na_2Cr_2O_7}$ is	hygroscopic in nature and	d is more sol	uble in water than ${ m K}_2{ m Cr}_2$	0 ₇ .	
	Therefore, o	ption C is correct.				



Q.21. Match the following

Column I	Column II
(i). Cotton mills	(a). Biodegradable waste
(ii). Paper mills	(b). Gypsum
(iii). Fertilizers	(c). Non-biodegradable waste
(iv). Thermal power plant	(d). Fly ash

- A) i c, ii a, b; iii c, iv bB) i - a; ii - a, iii - b, iv - d
- $\label{eq:constraint} \textbf{C} \qquad i-a,c; ii-b, iii-b, iv-a \qquad \qquad \textbf{D} \qquad i-c, ii-b,c; iii-b,c, iv-a$
- Answer: i a; ii a, iii b, iv d
- Solution: Cotton mills generate biodegradable waste such as cotton fiber, cottonseed, and other organic materials. Paper mills produce biodegradable waste such as pulp sludge, bark, and wood chips. Fertilizers are made from raw materials such as phosphate rock, sulfur, and potassium salts. During their production, gypsum is generated as a waste product. Thermal power plants generate non-biodegradable waste such as fly ash, bottom ash, and slag as by-products of coal combustion.
- Q.22. Match the column I with column II

Column I	Column II
A. Dacron	P. Thermosetting
B. Urea and formaldehyde resin	Q. Biodegradable
C. Nylon-2, Nylon-6	R. Polyester
D. Nylon-6,6	S. Uses for making bristles of brushes

 $\mathsf{A}) \qquad \mathsf{A}-\mathsf{R}, \mathsf{B}-\mathsf{P}, \mathsf{C}-\mathsf{S}, \mathsf{D}-\mathsf{Q}\mathsf{B}) \qquad \mathsf{A}-\mathsf{P}, \mathsf{B}-\mathsf{R}, \mathsf{C}-\mathsf{Q}, \mathsf{D}-\mathsf{S}\mathsf{C}) \qquad \mathsf{A}-\mathsf{R}, \mathsf{B}-\mathsf{P}, \mathsf{C}-\mathsf{Q}, \mathsf{D}-\mathsf{S}\mathsf{D}) \qquad \mathsf{A}-\mathsf{R}, \mathsf{B}-\mathsf{R}, \mathsf{C}-\mathsf{S}, \mathsf{D}-\mathsf{Q}$

Answer: A - R, B - P, C - Q, D - S

Solution: Dacron is a type of polyester fabric made from polyethylene terephthalate (PET). It is a thermoplastic polymer that is commonly used in clothing, bedding, and as a fibrefill material in various products.

Urea formaldehyde resin is a type of thermosetting plastic that is commonly used in the manufacture of adhesives, coatings, and molded products.

Nylon-2 and Nylon-6 are types of biodegradable nylon that can break down naturally in the environment over time. These types of nylon are typically used in applications such as packaging, textiles, and medical products.

Nylon-6, 6 is a type of nylon that is known for its high durability and abrasion resistance. It is commonly used in the manufacture of bristles for brushes, as well as in other applications such as automotive parts, fishing line, and clothing.

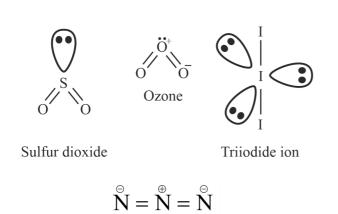
Q.23. How many of the following are bent in shape?

 SO_2, O_3, I_3^-, N_3^- .

2



- Solution: SO_2 has a bent molecular geometry due to the presence of a lone pair of electrons on sulfur atom, which distorts the molecular shape. In O_3 central oxygen atom is surrounded by two unshared electrons (one lone pair) and six bonding electrons (three bonds). So it is bent in shape.
 - I_3^- and N_3^- are linear.



Nitride ion

Q.24. The sum of number of lone pairs in central atom in IF_5 and IF_7 is:

Answer:

1

Solution: In IF_5 , the central atom is iodine, which has 7 valence electrons. Iodine in this molecule has 5 bonding pairs and 1 lone pair. Therefore, the number of lone pairs on the central atom of IF_5 is 1.

In IF₇, the central atom is also iodine, which has 7 valence electrons. Iodine in this molecule has 7 bond pairs and 0 lone pairs. Therefore, the number of lone pairs on the central atom of IF₇ is 0.

Therefore, the sum of the number of lone pairs in the central atom of IF₅ and IF₇ is 1 + 0 = 1.

Q.25. The degree of dissociation of monobasic acid is 0.3, By what percent is the observed depression in freezing point greater than the calculated depression in freezing point?

Answer: 30

Solution:

 $\overset{\text{HA}}{1-\alpha} \stackrel{\text{H}^+}{=} \overset{\text{H}^+}{\alpha} \overset{\text{A}^-}{+} \overset{\text{A}^-}{\alpha}$

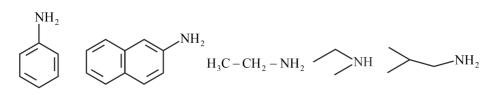
Van't Hoff factor, $i=1-\alpha+\alpha+\alpha=1+\alpha$

Given, degree of dissociation, $\alpha=0.\,3$

Therefore, i = 1 + 0.3 = 1.3

The observed depression in freezing point is 1.3 times of the calculated depression in freezing point. Therefore, the difference is 0.3 and the percentage is 30.

Q.26. How many compounds can be easily prepared by Gabriel Phthalimide synthesis, which on reaction with Hinsberg reagent produces a compound which is soluble in KOH?

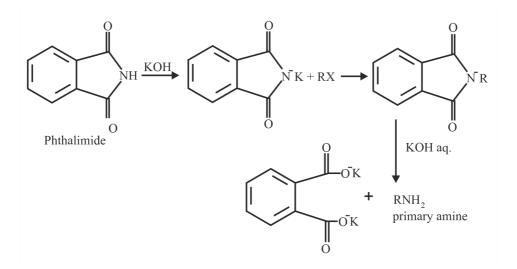


Answer:

2



Solution: Gabriel's Phthalimide synthesis is used for the conversion of primary alkyl halide into a primary amine. In Gabriel phthalimide synthesis, a base abstract proton from phthalimide gives a nucleophile phthalimide ion which attacks on the unhindered primary alkyl halide. The base hydrolysis of alkylated phthalimide gives the primary unhindered amine and phthalimide ion.



Mathematics

Q.27.	The negation of the sta	atement $(p \lor q) \land extsf{-}r$ is		
A) (~	$(p \wedge \neg q) \wedge r$	$B) \qquad ({\scriptstyle{{}^{\sim}}} p \wedge {\scriptstyle{{}^{\sim}}} q) \vee r$	$C) \qquad ({}^{\scriptscriptstyle \sim} p \wedge q) \vee r$	$D) \qquad (p \wedge {\scriptstyle{\sim}} q) \vee r$
Answer	$(\ensuremath{\sc r} p \wedge \ensuremath{\sc - q}) ee r$			
Solution	The negation of the	statement $(p \lor q) \land extsf{-}r$ is		
	~ $[(p \lor q) \land extsf{-}r]$			
	$\equiv \text{ ~} (p \lor q) \lor r$			
	$[\because \neg (A \land B) \equiv \neg A \lor \neg$	~B]		
	$\equiv (\texttt{-}p \land \texttt{-}q) \lor r$			
	$[\because \neg (A \lor B) \equiv \neg A \land \neg$	~B]		
Q.28.	From a square of side a cuboid with open top is		${ m cm}$ is cut off to make cuboid c	f maximum volume. The surface area of

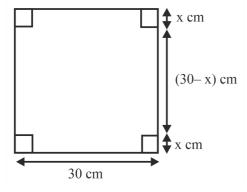
A)	400 cm^2	B)	464 cm^2	C)	800 cm^2	D)	$900 \ \mathrm{cm}^2$
	0						

Answer: 800 cm^2



Solution: Given that, the side of square is 30 cm and x cm squares are cut off.

The required diagram is



Now the dimensions of the cuboid formed will be

l(x) = 30 - 2x, b(x) = 30 - 2x and h(x) = x.

The Volume of the cuboid will be $V(x) = (30 - 2x)^2(x)$

Now to get Maximum value,

$$\Rightarrow \frac{\mathrm{d} V(x)}{\mathrm{d} x} = 0$$

$$\Rightarrow 2 (30 - 2x) (-2)x + (30 - 2x)^2 (1) = 0$$

$$\Rightarrow (30-2x)\left(-4x+30-2x\right)=0$$

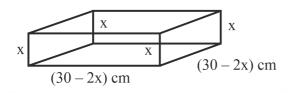
On simplifying we get,

 $\Rightarrow x = 15 \ cm, \ 5 \ cm$

But x cannot be $15\,\,{\rm cm}$ as the volume becomes zero.

Hence x = 5 cm.

Now to find the surface area of the cuboid,



Surface area will be $= (30 - 2x) \times x \times 4 + (30 - 2x)^2$

$$=(30-2 imes 5) imes 5 imes 4+(30-2 imes 5)^2$$

 $= 800 \text{ cm}^2.$

Therefore, the required surface area will be 800 cm^2 .

Q.29. Using the number 1, 2, 3, ..., 7, total numbers of 7 digit number which does not contain string 154 or 2367 is, (repetition is not allowed)

A)	4897	B)	4898	C)	4896	D)	4899
----	------	----	------	----	------	----	------



Solution:	Given,										
	The number $1, 2, 3, \ldots, 7$										
	Now total numbers of 7 digit number will be 7! without repetition,										
	Now total numbers which contain string 154 will be 5!										
	Total numbers which contain string 2367 will be 4!										
	And number which contain both string 2367 $\&~154$ will be 2										
	So, total number which does no	t contain string 154 or 2365	7 will be $7! - (5! + 4! - 2)$								
	=5040-(120+24-2)										
	=5040 - 142 = 4898										
Q.30. F	Find the value of $96 \cdot \cos \frac{\pi}{33} \cdot \cos \frac{2\pi}{33}$	$\cdot \cos \frac{4\pi}{33} \dots \cos \frac{16\pi}{33}$									
A) 0	B) 1	C)	2	D)	3						
Answer:	3										
Solution:	Given,										
	Expression $96 \cdot \cos \frac{\pi}{33} \cdot \cos \frac{2\pi}{33} \cdot \cos \frac{2\pi}{33}$	$\cos\frac{4\pi}{33}\dots\dots\cos\frac{16\pi}{33}$									
	Now we know that,										
	$\cos A \cdot \cos 2A \cdot \cos 2^2A \cdot \cos 2^3A.$	$\dots \cos 2^{n-1}A = \frac{\sin 2^n A}{2^n \sin A}$									
	Now using the above formula in	given expression we get,									
	$96 \cdot \cos\frac{\pi}{33} \cdot \cos\frac{2\pi}{33} \cdot \cos\frac{4\pi}{33} \dots$	$\ldots \cos \frac{16\pi}{33}$									
	$=96\times\frac{\displaystyle\frac{\sin\frac{32\pi}{33}}{2^{5}\displaystyle\frac{\pi}{33}}}{2^{5}\displaystyle\frac{\pi}{33}}$										
	$=96 imesrac{\sin{\left(\pi-rac{\pi}{33} ight)}}{2^{5}\sin{rac{\pi}{33}}}$										
	$=96 imes rac{\sin \left(rac{\pi }{33} ight) }{2^5 {\sin rac{\pi }{33}}} ~~ \{ { m as}~~ \sin \left(\pi -lpha ight)$	$=\sinlpha\}$									
	$=96\times\frac{1}{32}=3$										
Q.31.	Slope of tangent to a curve at a va	riable point is $rac{x^2+y^2}{2xy}$ and $rac{x^2}{2}$	y(2) = 0, then $y(8)$ is								
A) $\sqrt{2}$	B) $\sqrt{5}$	C)	$4\sqrt{3}$	D)	None of these						

Answer: $4\sqrt{3}$



Solution: Given:

$$\frac{dy}{dx} = \frac{x^2 + y^2}{2xy}$$
Put $y = vx \Rightarrow \frac{dy}{dx} = v + x\frac{dv}{dx}$
 $v + x\frac{dv}{dx} = \frac{1 + v^2}{2v}$
 $\Rightarrow x\frac{dv}{dx} = \frac{1 - v^2}{2v}$
 $\Rightarrow \int \left(\frac{2v}{v^2 - 1}\right) dv = -\int \frac{dx}{x}$
 $\Rightarrow \log_e |v^2 - 1| = \log_e \left(\frac{C}{x}\right)$
 $\Rightarrow \frac{y^2 - x^2}{x^2} = \frac{C}{x}$
 $\Rightarrow y^2 - x^2 = Cx$
Put $x = 2$ and $y = 0$ we get,
 $0 - 2^2 = 2C \Rightarrow C = -2$
 $\Rightarrow y^2 = x^2 - 2x$
 $\Rightarrow y(8) = \sqrt{8^2 - 16}$
 $\Rightarrow y(8) = \sqrt{48} = 4\sqrt{3}$

Q.32. If the order of the matrix A is 3×3 and |A| = 2, then the value of $\left| 3adj \left(|3A|A^2 \right) \right|$ is

A) $2^{21} \cdot 3^{10}$ B) $2^{10} \cdot 3^{21}$ C) $2^{12} \cdot 3^{15}$ D) $2^{15} \cdot 3^{12}$

Answer: $2^{10} \cdot 3^{21}$

Solution: We need to find the value of $\left|3adj\left(|3A|A^2\right)\right|$

We know that $|kA| = k^n |A|$ where *n* is the order of the matrix and *k* is a constant.

$$\Rightarrow |3A| = 3^{3}(2)$$

$$\Rightarrow |3adj (|3A|A^{2})| = 3^{3} |adj ((3^{3} \cdot 2)A^{2})|$$
We know that $adj (kA) = k^{n-1}adj (A)$

$$= 3^{3} |(3^{3} \cdot 2)^{2}adj (A^{2})|$$
Now we know that $|adjA| = |A|^{n-1}$

$$= 3^{3} ((3^{3} \cdot 2)^{2})^{3} |adj (A^{2})|$$

$$= 3^{3} (3^{3} \cdot 2)^{6} |A^{2}|^{3-1}$$

= $3^{3} (3^{3} \cdot 2)^{6} (2)^{4}$
= $2^{10} \cdot 3^{21}$

Hence, the required answer is $2^{10}\cdot 3^{21}$

Q.33. Find the number of integral values of x which satisfy the inequality $x^2 - 10x + 19 < 6$

A) 7		B)	10	C)	6	D)	8
Answer:	7						



Solution: Given: $x^2 - 10x + 19 < 6$ $\Rightarrow x^2 - 10x + 13 < 0$ Now, $x^2 - 10x + 13 = 0$ $\Rightarrow x = \frac{10 \pm \sqrt{48}}{2}$ $\Rightarrow x = 5 \pm 2\sqrt{3}$ So. $x^2 - 10x + 13 < 0$ $\Rightarrow \left(x-\left(5+2\sqrt{3}
ight)
ight)\left(x-\left(5-2\sqrt{3}
ight)
ight)<0$ $\Rightarrow x \in \left(5 - 2\sqrt{3}, 5 + 2\sqrt{3}\right)$ $\Rightarrow x \in (1.5, 8.4)$ So, integral values are x = 2, 3, 4, 5, 6, 7, 8 i.e., 7 values. The coefficient of x^7 in $\left(1-2x+x^3
ight)^{10}$ is Q.34. B) A) 54102080 C) 4080 D) 6234 Answer: 4080 We need to find the coefficient of x^7 in $\left(1-2x+x^3
ight)^{10}$. Solution: We know that for $(x + y + z)^n$, $T_n = rac{n!}{a^{|b|c|}} (x)^a (y)^b (z)^c$ such that a+b+c=nNow for $\left(1-2x+x^3
ight)^{10}$ $T_n = \frac{10!}{a!b!c!} (1)^a (-2x)^b \left(x^3\right)^c$ $=\frac{10!}{a!b!c!}-2^bx^{b+3c}$. with a+b+c=10Here we need coefficient of x^7 . Hence the combinations would be which satisfies both a + b + c = 10 and b + 3c = 7. Hence, coefficient of x^7 is $=\frac{10!}{3!7!0!}(-2)^7 + \frac{10!}{5!4!1!}(-2)^4 + \frac{10!}{7!1!2!}(-2)^1$ $= 120 \times (-128) + 20160 + (-720)$ =4080Therefore, the required value is 4080 If $a^2 + {(ar)}^2 + {\left({ar^2} \right)}^2 = 33033, (a,r \in N)$ then the value of $a + ar + ar^2$ is Q.35.

C)

230

B)

249

A)

Answer:

148

231

D)

231



Given that $a^2 + (ar)^2 + (ar^2)^2 = 33033, (a, r \in N)$ $\Rightarrow a^2 \left(1+r^2+r^4
ight) = 11^2 imes 273$ On comparing both sides of the equation we get, a = 11 and $1 + r^2 + r^4 = 273$ $\Rightarrow r^2 + r^4 = 272$ $\Rightarrow r^2 \left(1+r^2
ight) = 16 imes 17$ $\Rightarrow r = 4$ Now $a + ar + ar^2 = 11 + 11 \times 4 + 11 \times 16$ = 11 + 44 + 176 = 231.Therefore, the required value is 231 Q.36. Shortest distance between the lines $\frac{x+1}{7} = \frac{y+1}{-6} = \frac{z+1}{1}$ and $\frac{x-3}{1} = \frac{y-5}{-2} = \frac{z-7}{1}$ is $2\sqrt{29}$ units C) $3\sqrt{29}$ units B) A) $\sqrt{29}$ units $2\sqrt{29}$ units Answer: Solution: Given: $\frac{x+1}{7} = \frac{y+1}{-6} = \frac{z+1}{1}$ $\frac{x-3}{1} = \frac{y-5}{-2} = \frac{z-7}{1}$ So. $\overrightarrow{a}_1 = -\hat{\imath} - \hat{\jmath} - \hat{k}$ $\overrightarrow{a}_2 = 3\hat{i} + 5\hat{j} + 7\hat{k}$ So. $\overrightarrow{a}_2 - \overrightarrow{a}_1 = 4\hat{i} + 6\hat{j} + 8\hat{k}$ And. $\overrightarrow{b}_1 \times \overrightarrow{b}_2 = \begin{vmatrix} \hat{\imath} & \hat{\jmath} & \hat{k} \\ 7 & -6 & 1 \\ 1 & -2 & 1 \end{vmatrix} = -4\hat{\imath} - 6\hat{\jmath} - 8\hat{k}$ $\Rightarrow \left|\overrightarrow{b}_{1} \times \overrightarrow{b}_{2}\right| = \sqrt{16 + 36 + 64} = \sqrt{116}$ Shortest distance between the lines $= \frac{\left|\frac{\left(\overrightarrow{a}_{2} - \overrightarrow{a}_{1}\right) \cdot \left(\overrightarrow{b}_{1} \times \overrightarrow{b}_{2}\right)}{\left|\overrightarrow{b}_{1} \times \overrightarrow{b}_{2}\right|}\right|}$ $= \left|\frac{\left(4\hat{\imath}+6\hat{\jmath}+8\hat{k}\right)\cdot\left(-4\hat{\imath}-6\hat{\jmath}-8\hat{k}\right)}{\sqrt{116}}\right|$ $=\left|\frac{-16-36-64}{\sqrt{116}}\right|=\sqrt{116}=2\sqrt{29}$ units Q.37. If 3, 8, 13, 373 is in AP then sum of terms which is not divisible by 3 is

9525 Answer:

Solution:

D)

 $5\sqrt{29}$ units



Solution: Given sequence is $3, 8, 13, \dots, 373$.

Here a = 3, d = 5 and $a_n = 373$

Now $a_n = a + (n-1)d$

 $\Rightarrow 373 = 3 + (n-1)5$

 $\Rightarrow n = 75$

We know that $S_n = \frac{n}{2}(a + a_n)$

$$\Rightarrow S_{75} = \frac{75}{2} \left(3 + 373\right)$$

 $\Rightarrow S_{75} = 14100$

Now let us write the sequence of terms which are divisible by 3.

We get, 3, 18, 33, 363

$$\Rightarrow 363 = 3 + (n-1)15$$

 $\Rightarrow n = 25$

Now let us find the sum of terms divisible by 3.

$$\Rightarrow S_{div\;by\;3} = \frac{25}{2} \left(3 + 363\right) = 4575$$

Required sum = Sum of 75 terms - Sum of terms divisible by 3

= 14100 - 4575

= 9525.

Therefore, the required sum is 9525

Q.38. If the coefficient of x^7 in the expansion of $\left(ax - \frac{1}{bx^2}\right)^{13}$ is equal to the coefficient of x^{-5} in the expansion of $\left(ax + \frac{1}{bx^2}\right)^{13}$, then a^4b^4 is



Solution: The coefficient of x^7 in the expansion of $\left(ax^2 - \frac{1}{bx}\right)^{13}$ is equal to the coefficient of x^{-5} in $\left(ax + \frac{1}{bx^2}\right)^{13}$.

We know that, the general term T_{r+1} in the expansion $\left(\mathrm{a}+\mathrm{b}\right)^n~$ is

$$\mathrm{T}_{r+1} = {}^{n} \mathrm{C}_{r} \mathrm{a}^{n-r} \mathrm{b}^{r}$$

Applying to
$$\left(ax - \frac{1}{bx^2}\right)^{13}$$
, we get
 $T_{r+1} = {}^{13}C_r(ax){}^{13-r}\left(-\frac{1}{bx^2}\right)^r$
 $\Rightarrow T_{r+1} = (-1)^r \times {}^{13}C_r(a){}^{13-r}(x){}^{13-3r}(b){}^{-r}$

Therefore, $13 - 3r = 7 \Rightarrow r = 2$ for coefficient of x^7 .

Thus,

$$T_3 = {}^{13}C_2 \left(\frac{a^{11}}{b^2}\right)$$

Similarly, applying to $\left(ax+rac{1}{bx^2}
ight)^{13}$, we get

$$T_{r+1} = {}^{13}C_r(ax){}^{13-r}\left(\frac{1}{bx^2}\right)^r$$
$$\Rightarrow T_{r+1} = {}^{13}C_r(a){}^{13-r}(x){}^{13-3r}(b){}^{-r}$$

Therefore, 13 - 3r = -5 for coefficient of x^{-5}

 $\Rightarrow r = 6$

So,

$$T_7 = {}^{13}C_6(a)^7(b)^{-6}$$

Hence, applying the given condition we get

$$\begin{split} ^{13}C_2 \left(\frac{a^{11}}{b^2}\right) &= {}^{13}C_6(a)^7(b)^{-6} \\ \Rightarrow a^4b^4 &= \frac{13}{13}C_6 \\ \Rightarrow a^4b^4 &= \frac{13!}{7!\cdot 6!} \times \frac{2!\cdot 11!}{13!} \\ \Rightarrow a^4b^4 &= \frac{11\times 10\times 9\times 8}{6\times 5\times 4\times 3} \\ \Rightarrow a^4b^4 &= 22 \end{split}$$

Q.39. Two dice are rolled and sum of two digits is N then probability that $2^N < N!$ is $\frac{m}{n}$, where m and n are coprime, then 11m - 3n is





N is the sum of numbers of two dice.

 $\Rightarrow 2 \leq N \leq 12$

Let us check the given condition $2^N < N!$

This condition is only true for $N \ge 4$

We need to find the probability for $N \ge 4$.

$$\Rightarrow P(N \ge 4) = 1 - P(N = 2) - P(N = 3)$$

$$= 1 - \frac{1}{36} - \frac{1}{36}$$
$$= \frac{11}{12} = \frac{m}{n}$$

Now let us find 11m - 3n.

$$= 11 \times 11 - 3 \times 12 = 85.$$

Hence, the requried answer is 85.